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C cc= 6 protons + 2

electrons= 4. What

trend in atomic

radius is observed as

one moves down a

group in the periodic

table? The radius

increases because the

valence electrons are

getting further away

from the nucleus.

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Atoms and ions have
six protons

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Isotope Atomic

Number Z Mass

Number A Number of

Electrons 27 59 25

^{14}N 3 7 3 3 6 3

$^{58}\text{Zn}^{2+}$ $^{19}\text{F}^{-}$ 10.

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sentences, describe what the isotopes of an element have in common and how they are different. 11.

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because they are in

the same valence

shell and Kr has the

higher core charge

(+8 vs. +7). IE of Rb is

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the lowest because core charge is +1 and its valence shell ($n = 5$) is larger than the valence shell ($n = 4$) of Kr and Br. 7. One of the inner shell electrons is harder to remove because it is closer to the nucleus.

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4. All carbon atoms
and ions have six
protons in the
nucleus. 5. All
hydrogen atoms and
ions have one proton
in the nucleus. 6.

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